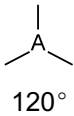
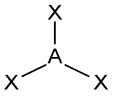
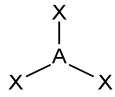
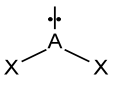
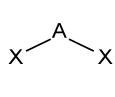

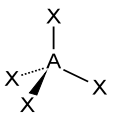
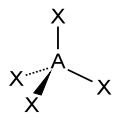
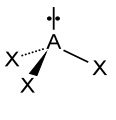
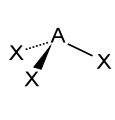
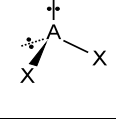
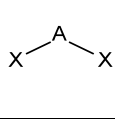
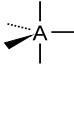
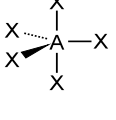
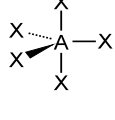
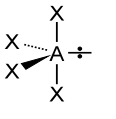
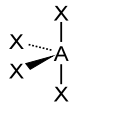

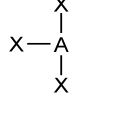
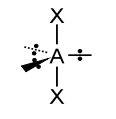

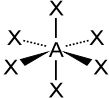
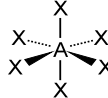
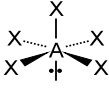
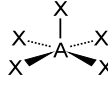
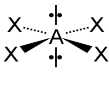
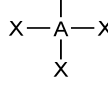
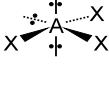
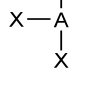
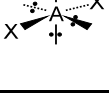
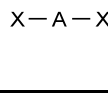


Molecular Shapes Predicted by VSEPR Theory

| Number of electron groups around central atom | Arrangement of electron groups | Number of atoms bonded to the central atom | Number of lone pairs on central atom | VSEPR notation | Electron geometry | Molecular geometry (Shape diagram) | Shape name |
|---|--|--|--------------------------------------|----------------|--|---|----------------------|
| 2 | —A— 180° | 2 | 0 | AX_2 | X—A—X | X—A—X | linear |
| 3 |  120° | 3 | 0 | AX_3 |  |  | trigonal planar |
| | | 2 | 1 | AX_2E |  |  | bent |
| 4 |  109.5° | 4 | 0 | AX_4 |  |  | tetrahedral |
| | | 3 | 1 | AX_3E |  |  | trigonal pyramidal |
| | | 2 | 2 | AX_2E_2 |  |  | bent |
| 5 |  90°, 120° | 5 | 0 | AX_5 |  |  | trigonal bipyramidal |
| | | 4 | 1 | AX_4E |  |  | seesaw |
| | | 3 | 2 | AX_3E_2 |  |  | T-shaped |
| | | 2 | 3 | AX_2E_3 |  | X—A—X | linear |

| Number of electron groups around central atom | Arrangement of electron groups | Number of atoms bonded to the central atom | Number of lone pairs on central atom | VSEPR notation | Molecular geometry | Shape diagram | Shape name |
|---|--|--|--------------------------------------|----------------|---|--|------------------|
| 6 |  90° | 6 | 0 | AX_6 |  |  | octahedral |
| | | 5 | 1 | AX_5E |  |  | square pyramidal |
| | | 4 | 2 | AX_4E_2 |  |  | square planar |
| | | 3 | 3 | AX_3E_3 |  |  | T-shaped |
| | | 2 | 4 | AX_2E_4 |  |  | linear |